

Rates of Reaction Experiment

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Rates of Reaction Experiment

INTRODUCTION

This experiment is called 'rates of reaction', I will be partaking in this experiment to find out what factors effect the 'rate of reaction' between magnesium (strips) and hydrochloric acid. Each time

I repeat the experiment, I will be adding 0.5 moles of hydrochloric acid and recording the volume of gas given out during the testing and record the time it takes to give the reaction needed. A reaction is taken place within the [experiment](#) because there is a [collision](#) of particles in each element. When these two elements collide with each other they form a chemical reaction (a change which is irreversible) we know this because a gas is given of too show a change, this is why we record the amount of gas given off to show when the reactions takes place and how big of a change it has made.

WORD EQUATION:

Magnesium + Hydrochloric Acid → Magnesium Chloride + Hydrogen Acid

CHEMICAL EQUATION



Activation energy is the minimum amount of energy, which the particles must have in order to react.

There are five factors, which can make a difference to the rate of a reaction. *1* The surface of the element

2 Catalyst, the amount of the substance

3 The concentration of the element

4 Temperature, due the heat of the atmosphere

5 Pressure pronounced to the substance

When a reaction takes place, the substances used are called reactants and the substances produced are called products. Because the amount of products being produced are increasing the number of reactants are or should be decreasing.

To calculate if and how well products are increasing and reactants are decreasing we use a number of ways, we firstly find out how much of the reactant(s) is being used up and how much product(s) we are gaining from this in a set time.

FORMULARS TO SOLVE THIS:

*** How much reactant is being used up***

Amount of reactant being used up

Reaction rate = Time taken

How much product's being produced in a set time

Amount of product formed

Reaction rate = Time taken

A reaction can only take place when a successful collision is occurred, so for a reaction to take place two different elements have to concur with each other, this depends on the amount of atoms and energy in the two elements, this is called a successful collision. A chemical reaction can only take place when two different elements collide with each other.

There are other methods to measure the rate (speed) of reaction.

Precipitation: observe a marker through the solution and predict how long it takes to vanish.

A change in mass: a mass balance can only be carried out during the

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